

CH105 SI  
Session 4  
Bliss Chang

1. Name 3 chemical species isoelectronic with Ne.



2. Name the following ions:

$Na^+$	sodium	$O^{2-}$	oxide
$F^-$	fluoride	$Co^{2+}$	cobalt(II)
$Be^{2+}$	beryllium	$Ba^{2+}$	barium
$Fe^{3+}$	iron(II)	$Br^-$	bromide

3. Name the following ionic compounds:

$NaCl$	sodium chloride	$RbCl$	rubidium chloride
$KF$	potassium fluoride	$CuI_2$	copper(II) iodide
$CaI_2$	calcium iodide	$FeF_3$	iron(II) fluoride
$LiBr$	lithium bromide	$NaOH$	sodium hydroxide
$BaO$	barium oxide	$NH_4NO_3$	ammonium nitrate
$KCN$	potassium cyanide	$BaCO_3$	barium carbonate
$SrSO_4$	Strontium sulfate	$KHCO_3$	potassium bicarbonate
			potassium hydrogen carbonate

4. Provide chemical formulas for the following:

Zinc(II) chloride	$ZnCl_2$	Magnesium sulfate	$MgSO_4$
Lithium iodide	$LiI$	Nickel(II) chloride	$NiCl_2$
Cesium oxide	$Cs_2O$	Potassium chlorate	$KClO_4$

5. Draw the Lewis dot structure of the following atoms or ions:

$Ca$	$Ca:$	$O^{2-}$	$:O^-$
$Xe$	$:Xe:$	$B$	$:B$
$P$	$:P$	$Ba$	$:Ba$

6. Order the following from lowest to highest ionization energy:

